

Development and applications of the LFDFT: the non-empirical account of ligand field and the simulation of the f–d transitions by density functional theory†‡

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Ligand field density functional theory (LFDFT) is a methodology consisting of non-standard handling of DFT calculations and post-computation analysis, emulating the ligand field parameters in a non-empirical way. Recently, the procedure was extended for two-open-shell systems, with relevance for inter-shell transitions in lanthanides, of utmost importance in understanding the optical and magnetic properties of rare-earth materials. Here, we expand the model to the calculation of intensities of f → d transitions, enabling the simulation of spectral profiles. We focus on Eu²⁺-based systems: this lanthanide ion undergoes many dipole-allowed transitions from the initial 4f⁷(⁸S_{7/2}) state to the final 4f⁶5d¹ ones, considering the free ion and doped materials. The relativistic calculations showed a good agreement with experimental data for a gaseous Eu²⁺ ion, producing reliable Slater–Condon and spin–orbit coupling parameters. The Eu²⁺ ion-doped fluorite-type lattices, CaF₂:Eu²⁺ and SrCl₂:Eu²⁺, in sites with octahedral symmetry, are studied in detail. The related Slater–Condon and spin–orbit coupling parameters from the doped materials are compared to those for the free ion, revealing small changes for the 4f shell side and relatively important shifts for those associated with the 5d shell. The ligand field scheme, in Wybourne parameterization, shows a good agreement with the phenomenological interpretation of the experiment. The non-empirical computed parameters are used to calculate the energy and intensity of the 4f⁷–4f⁶5d¹ transitions, rendering a realistic convoluted spectrum.

Introduction

The concept of the ligand field, very fruitful in the effective account of bonding and properties in coordination chemistry, is equivalent to the crystal field theory in condensed matter science. Both terminologies refer to the same phenomenological model, operated with adjustable parameters.

Born more than eighty years ago, from the work of H. Bethe¹ and J. H. van Vleck² it still keeps the position of the most transparent way to describe the optical and magnetic properties of metal ion-based systems (lattices or molecular complexes). As long as quantum chemical methods can compute reliable energy level schemes, the subsequent ligand field analysis of the raw results is the way to illuminate in depth the underlying mechanism.^{3–5} *Stricto sensu*, the ligand field refers to effective one-electron parameters accounting for the effect of the environment on a metal ion, but the complete frame includes the inter-electron effects, describing the electronic correlation in the active space of dⁿ or fⁿ configurations, and also the spin–orbit coupling, namely the relativistic effects. Besides the standard theory, one must note the paradigm shift due to C. E. Schäffer and C. K. Jørgensen, who revisited the ligand field theory to ensure more chemical insight within their Angular Overlap Model (AOM), initially devoted to the d-type transition metal systems.⁶ W. Urland pioneered this model for the f-type ligand field, in lanthanide compounds, with convincing applications in spectroscopy and magnetism.⁷

About two decades ago, given the important growth of computational techniques, the demand for a predictive theory compatible

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‡ Electronic supplementary information (ESI) available: The whole range of the spectral energy obtained for CaF₂:Eu²⁺ and SrCl₂:Eu²⁺ together with a direct comparison between the theoretical spectra and the excitation spectra in ref. 41 and 42.

with the classical formalism of the ligand field theory emerged. In particular, this is not a trivial task in the frame of density functional theory (DFT), limited to non-degenerate ground states, while ligand field concerns the full multiplets originating from d^n or f^n configurations. In the consistent solving of this problem, C. Daul *erat primus*. He and co-workers (noting the contribution of M. Atanasov) designed a pioneering approach by non-routine handling of DFT numeric experiments, to extract ligand field parameters, in a post-computational algorithm named LFDFT.^{8–10} The procedure treats the near degeneracy correlation explicitly within the model space of the Kohn–Sham orbitals possessing dominant d and f characters.

In LFDFT, the basic start is a DFT calculation performed in average of configuration (AOC) conditions. Namely, for a given d^n (or f^n) configuration of the metal ion in the complex, the occupation of five (or, respectively, seven) Kohn–Sham orbitals carrying main d (or f) character is fixed to the general fractional $n/5$ (or $n/7$, respectively) numbers. This corresponds to the barycentre conceived in formal ligand field theories. Subsequently, with the converged AOC orbitals, a series of numeric experiments are done, producing the configurations related to the distribution of n electrons in the five (or seven) orbitals identified as the ligand field sequence (this time with corresponding integer populations). These determinant configurations are not real states, but useful computational experiments, able to render ligand field parameters. The situation is somewhat similar to broken symmetry treatments,^{11–14} where the spin-polarized configurations cannot be claimed as physical states, but artificial constructions relevant for the emulation of the exchange coupling parameters.¹⁵ Then, the LFDFT run of different configurations based on AOC orbitals yields ligand field parameters, altogether with inter-electron Coulomb and exchange effective integrals. Thus, the Slater determinants are used as the basis in the computational model. In the advanced background of the theory, a canonical number of configurations needed to reproduce the desired parameters can be defined as a function of the symmetry of the problem (Slater determinant wavefunctions of spin-orbitals weighted by symmetry coefficients).¹⁰ In practice, the full set of configurations can be generated, performing the least square fit relating the computed energy expectation values against the ligand field model formulas. The obtained parameters are further used in setting configuration interaction (CI) matrices, in the spirit of the ligand field formalism, sustained in a non-empirical manner. Therefore C. Daul *et al.* have realized the parameter-free ligand field theory, which became a valuable tool for any consideration of multiplet states in DFT.

We recognize herein the impact of the LFDFT in solving various electronic structure problems. This computational gadget has revolutionized many fields of chemical science, being applied in theoretical investigations^{16–20} as well as in experimental works.^{21,22}

A priori, LFDFT has determined the multiplet energy levels within an accuracy of a few hundred wavenumbers.²³ The model has given satisfactory results for the molecular properties arising from a single-open-shell system, such as zero-field splitting (ZFS),^{24,25} magnetic exchange coupling,^{26–29} Zeeman interaction,³⁰ hyper-fine splitting,³⁰ shielding constants,^{31,32} d–d and f–f transitions.^{10,17,33,34}

Recently, the LFDFT algorithm has been updated to handle the electronic structure of two-open-shell systems, as it is important in the understanding of the optical manifestation of lanthanide phosphors.^{35,36} Lanthanide compounds are agents in light-emitting diode (LED) technology, used in domestic lighting.³⁷ In the case of a two-open-shell inter-configuration of f and d electrons, the size of the ligand field CI matrices is collected in Table 1, calculated with the following combinatorial formulas:

$$N(4f^n) = \binom{14}{n}, \quad (1)$$

$$N(4f^{n-1}5d^1) = \binom{14}{n-1} \cdot \binom{10}{1}, \quad (2)$$

as a function of the number of active electrons (n). We can confine to a single f–d orbital promotion, since the energy of two and further electron processes is too high. It is seen from Table 1 that the size of the CI matrices increases drastically, for some cases ($n = 7$ or 8) a parallelized algorithm having been required to achieve calculations.

In this paper, we present new development and applications of the LFDFT algorithm, previously validated for the two-open-shell $4f^15d^1$ electronic structure of Pr^{3+} .^{35,38–40} Special attention will be paid to Eu^{2+} systems, *i.e.* for $n = 7$ (Table 1), taking as examples divalent europium doped in the fluorite-type lattices CaF_2 and SrCl_2 , comparing the first principles results with the available experimental data.^{41,42}

Methodology

The two-open-shell ligand field-based CI Hamiltonian in eqn (3) combines quantum effects due to the inter-electron repulsion and exchange (H_{EE}), the spin-orbit coupling (H_{SO}) and the ligand field effective one-electron (H_{LF}):³⁵

$$H = H_0 + H_{\text{EE}} + H_{\text{SO}} + H_{\text{LF}}, \quad (3)$$

Table 1 Total number of generated Slater determinants corresponding to the $4f^n$ and $4f^{n-1}5d^1$ electron configurations of lanthanide ions with n valence electrons

n	1	2	3	4	5	6	7	8	9	10	11	12	13	14
$N(4f^n)$	14	91	364	1001	2002	3003	3432	3003	2002	1001	364	91	14	1
$N(4f^{n-1}5d^1)$	10	140	910	3640	10 010	20 020	30 030	34 320	30 030	20 020	10 010	3640	910	140
\sum^a	24	231	1274	4641	12 012	23 023	33 462	37 323	32 032	21 021	10 374	3731	924	141

^a \sum represents the cumulative sum of $N(4f^n)$ and $N(4f^{n-1}5d^1)$.

where H_0 is a diagonal matrix, which gathers contributions of zeroth-order interactions, such as the kinetic energy background and the nuclear–electron attraction of the AOC configuration:

$$H_0 = -\frac{\hbar^2}{2m} \sum_i \nabla_i^2 - \sum_i \frac{Ze^2}{r_i}. \quad (4)$$

This term acts only on the diagonal of the full ligand field CI matrix:

$$H_0 = \begin{pmatrix} I_{N(4f^n)} \cdot 0 & (0) \\ (0) & I_{N(4f^{n-1}5d^1)} \cdot \Delta(\text{fd}) \end{pmatrix}, \quad (5)$$

where I_N is an identity matrix of dimension N (see Table 1) and $\Delta(\text{fd})$ is the gap parameter, which determines the energy difference between the barycentre of the multiplet levels of the excited $4f^{n-1}5d^1$ and of those of the ground $4f^n$ electron configuration. Therefore, in further consideration, the H_{EE} , H_{SO} and H_{LF} matrices are simply traceless blocks, their possible diagonal elements already engulfed in $\Delta(\text{fd})$.

The matrix elements of H_{EE} are constructed from the two-electron integrals:

$$\langle \psi_a \psi_b | H_{EE} | \psi_c \psi_d \rangle = \int \psi_a^*(r_1) \psi_b^*(r_2) \frac{1}{r_{12}} \psi_c(r_1) \psi_d(r_2) dr_1 dr_2, \quad (6)$$

where ψ denotes the atomic orbital wavefunctions:

$$\psi(r) = R_{nl}(r) Y_{lm}(\theta, \phi), \quad (7)$$

R_{nl} is the radial wavefunction of the atomic shell and Y_{lm} is the spherical harmonic component. It is a basic assumption of the ligand field framework that the two-electron part can be treated like in the free atom.³⁵

Within mathematical operations, eqn (6) is reducible into the product of two integrals of angular and radial components. Once the angular part is explicitly resolved, the whole variety of the eqn (6) integrals can be represented by a few radial Slater–Condon parameters, F_k (eqn (8) and (9)) and G_k (eqn (10)), with intra- or inter-shell nature. In the two-open-shell problem of 4f and 5d electrons, one obtains:

$$F_k(\text{ff}) = \int_0^\infty \int_0^\infty \frac{r_<^k}{r_>^{k+1}} R_{4f}^2(r_1) R_{4f}^2(r_2) r_1^2 r_2^2 dr_1 dr_2, \quad (8)$$

$$F_k(\text{fd}) = \int_0^\infty \int_0^\infty \frac{r_<^k}{r_>^{k+1}} R_{4f}^2(r_1) R_{5d}^2(r_2) r_1^2 r_2^2 dr_1 dr_2, \quad (9)$$

$$G_k(\text{fd}) = \int_0^\infty \int_0^\infty \frac{r_<^k}{r_>^{k+1}} R_{4f}(r_1) R_{5d}(r_2) R_{5d}(r_1) R_{4f}(r_2) r_1^2 r_2^2 dr_1 dr_2. \quad (10)$$

The matrix elements of H_{SO} express the spin–orbit structure of the electronic multiplets. The formulation of H_{SO} has been the subject of numerous investigations^{3,43–45} where its matrix elements have been reasonably well approximated in atomic-like integrals:

$$\langle n l s m_l m_s | H_{SO} | n l s' m_l' m_s' \rangle = \zeta_{nl} \langle l s m_l m_s | \hat{l} \cdot \hat{s} | l s' m_l' m_s' \rangle, \quad (11)$$

where ζ_{nl} is the effective one-electron spin–orbit coupling constants for one electron in a nl atomic shell. It can be analytically evaluated using the radial wavefunction R_{nl} of the atomic shell:

$$\zeta_{nl} = \frac{Ze^2 \hbar^2}{8\pi\epsilon_0 m_0^2 c^2} \left\langle R_{nl} \left| \frac{1}{r^3} \right| R_{nl} \right\rangle. \quad (12)$$

The matrix elements of H_{LF} play the role of the chemical environment of the lanthanide ion. The general formulation of the ligand field potential follows Wybourne:⁴⁶

$$\begin{aligned} & \langle l_a m_{l_a} | H_{LF} | l_b m_{l_b} \rangle \\ &= \sum_{k=0}^{l_a+l_b} \sum_{q=-k}^k B_q^k(l_a, l_b) \langle Y_{l_a m_{l_a}}(\theta, \phi) | C_q^{(k)}(\theta, \phi) | Y_{l_b m_{l_b}}(\theta, \phi) \rangle, \end{aligned} \quad (13)$$

where $C_q^{(k)}$ represent the solid spherical harmonic tensor operators (eqn (11)) and B_q^k are the Wybourne-normalized crystal field parameters;

$$C_q^{(k)} = \sqrt{\frac{4\pi}{2k+1}} Y_{kq}. \quad (14)$$

The collection of non-vanishing Wybourne parameters depends on the coordination symmetry of the lanthanide centre, their total number in a two-open-shell f–d ligand field problem being 64 in the case of the C_1 point group.⁴⁷ Here, they cannot be reduced to having a simple electrostatic origin, since the DFT calculation takes into consideration different effects including orbital overlap and covalence.³⁵

Besides the Hamiltonian setting, other specific construction regards the matrix element of the dipole moment operator, important to the computation of the intensity of transitions:

$$\langle \psi_\mu | \vec{d}_z | \psi_\nu \rangle = \frac{1}{\sqrt{3}} \langle R_{n_\mu l_\mu} | r | R_{n_\nu l_\nu} \rangle \langle Y_{l_\mu m_\mu} | C_z^{(1)} | Y_{l_\nu m_\nu} \rangle, \quad (15)$$

where, in the right hand side of eqn (15), the term carrying the radial component is simple overlap integrals, while the angular term is proportional to Clebsch–Gordan coefficients.³⁹ Actually, only the f–d elements are non-vanishing, their mutual mixing by ligand field rendering the intensity, in an approximate, but apparently satisfactory manner.

In summary, several series of parameters have to be determined non-empirically in order to perform LFDFT calculations of two-open-shell f and d electrons:

(1) $\Delta(\text{fd})$, which represents the energy shift of the multiplets of the $4f^{n-1}5d^1$ configuration with respect to those of the $4f^n$ configuration.

(2) $F_k(\text{ff})$, $F_k(\text{fd})$ and $G_k(\text{fd})$, which represent the static electron correlation within the $4f^n$ and $4f^{n-1}5d^1$ configurations.

(3) ζ_{nl} , which represents the relativistic spin–orbit interaction in the 4f and 5d shells.

(4) $B_q^k(\text{f,f})$, $B_q^k(\text{d,d})$ and $B_q^k(\text{f,d})$, which describe the interaction due to the presence of the ligands onto the electrons of the metal centre.

The DFT calculations have been carried out by means of the Amsterdam density functional (ADF) program package (ADF2013.01).^{48–50} We must point out that the ADF is one of

the few DFT codes that has the set of keywords facilitating the AOC calculations and Slater determinant emulation, needed by the LFDFT procedure.^{35,36} The hybrid B3LYP functional⁵¹ was used to compute the electronic structure and the related optical properties, in line with previous works.^{35,36,39} The molecular orbitals were expanded using triple-zeta plus two polarization Slater-type orbital (STO) functions (TZ2P+) for the Eu atom and triple-zeta plus one polarization STO function (TZP) for the Ca, Sr, F and Cl atoms.

The geometrical structures due to the doping of the Eu^{2+} ion into CaF_2 and SrCl_2 lattices were approached *via* periodical calculations by means of the VASP program package.⁵² The local density approximation (LDA) defined in the Vosko-Wilk-Nusair (VWN)⁵³ and the generalized gradient approximation (GGA) outlined in the Perdew-Burke-Ernzerhof (PBE)⁵⁴ were used for the exchange–correlation functional. The interaction between valence and core electrons was emulated with the projected augmented wave method.^{55,56} External as well as semi-core states were included in the valence. A plane-waves basis set with a cut-off energy of 400 eV was used. Super-cells representing a 2 by 2 by 2 expansion of the unit cells of CaF_2 and SrCl_2 were simulated, which were found to be large enough to lead to negligible interactions between the periodic images of the Eu^{2+} impurity. 4 k -points were included in each direction of the lattice. The atomic positions were allowed to relax until all forces were smaller than $0.005 \text{ eV } \text{\AA}^{-1}$.

Results and discussion

The determination of $\Delta(\text{fd})$

We must discuss at the very beginning the calculation of the $\Delta(\text{fd})$ gap, which is important in the problem of two-open-shell systems because it sets the origin of the energy of the two electron configurations, conventionally 0 for the $4f^n$ and $\Delta(\text{fd})$ for the $4f^{n-1}5d^1$. In a first description, we work with the free ion, considering the gaseous Eu^{2+} case. Estimating by DFT the energy difference between the $4f^6 5d^1$ and $4f^7$ configurations, we must work under the AOC references. This is because the $\Delta(\text{fd})$ is not the difference between the specific energy levels, but rather a gap of the averaged energy values common for all multiplets of $4f^n$ and $4f^{n-1}5d^1$ kinds.

T. Ziegler *et al.* clarified early that the occupation-averaged configurations, called transition states, carry in DFT the meaning of statistically-averaged spectral terms.⁵⁷ We prepare the wavefunctions ψ_{4f} and ψ_{5d} by AOC where six and one electrons are evenly distributed in the 4f and 5d orbitals of Eu^{2+} , respectively (Fig. 1). This will generate the reference totally symmetric density, which will be used to compute the DFT energy associated with the series of Slater determinants. Thus all the Slater determinant energies are successively computed permuting seven electrons in the 4f wavefunction (Fig. 1) for the $4f^7$ manifold, and permuting six electrons in the 4f wavefunction plus one electron in the 5d for the $4f^6 5d^1$ manifold. The results obtained at the B3LYP level of theory are graphically represented in Fig. 2 showing the $\Delta(\text{fd})$ gap. Note that $\Delta(\text{fd})$ can

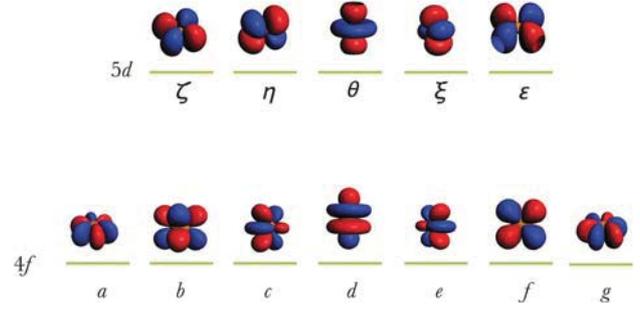


Fig. 1 Representation of the ψ_{4f} and ψ_{5d} orbitals of Eu^{2+} , obtained from an AOC calculation of Eu^{2+} within the $4f^6 5d^1$ electron configuration. The components of the 4f orbitals are listed from left to right according to: $f_{x(x^2-3y^2)}$, f_{xyz} , f_{z^2x} , f_{z^2y} , $f_{z(x^2-y^2)}$ and $f_{y(3x^2-y^2)}$, i.e. a, b, c, d, e, f and g. The components of the 5d orbitals are listed from left to right according to: d_{xy} , d_{xz} , d_{yz} , d_{z^2} and $d_{x^2-y^2}$, i.e. ζ , η , θ , ξ and ϵ .

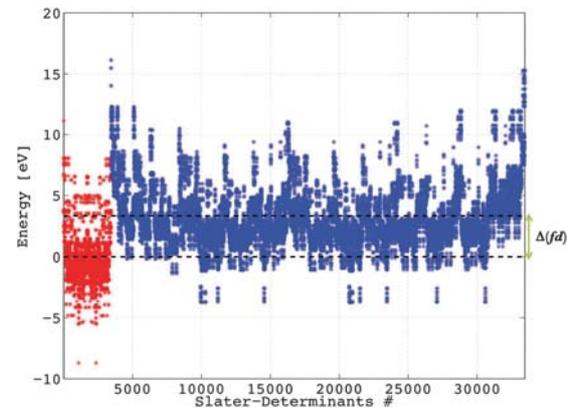


Fig. 2 Representation of the calculated DFT energy values associated with the 3432 Slater determinants (in red) arising from the $4f^7$ and the 30 030 Slater determinants (in blue) arising from the $4f^6 5d^1$ configurations of Eu^{2+} . The two dashed lines represent the barycentre of the $4f^7$ manifold (set to the zero of energy), and that of the $4f^6 5d^1$ manifold.

occasionally have a negative value, indicating that the ground electron configuration of the lanthanide ion is $4f^{n-1}5d^1$ instead of $4f^n$. Such a situation may appear in the case of lanthanide Gd^{2+} ($n = 8$, see Table 1) and La^{2+} ($n = 1$, see Table 1) ions.

The lowest energies corresponding to the $4f^7$ manifold (Fig. 2) are associated with the Slater determinants:

$$|a^+ b^+ c^+ d^+ e^+ f^+ g^+| \quad \text{and} \quad |a^- b^- c^- d^- e^- f^- g^-|,$$

where the signs + and – represent the spin of one electron, up or down, respectively. The highest energies corresponding to the $4f^6 5d^1$ manifold (Fig. 2) are associated with the Slater determinants:

$$|c^\pm d^\pm e^\pm \theta^+| \quad \text{and} \quad |c^\pm d^\pm e^\pm \theta^-|,$$

where the sign \pm represents a restricted occupation of two electrons in one orbital. The calculated value of the $\Delta(\text{fd})$ parameter is 3.10 eV at the B3LYP level of theory.

The DFT Slater determinant energies (Fig. 2) can also provide information about the two-electron $F_k(\text{ff})$, $F_k(\text{fd})$ and $G_k(\text{fd})$ parameters using Slater’s rule³ and least mean square fitting.¹⁰

However, this procedure might undergo uncertainty caused by the important number of linear equations *versus* variables. In the case of two-open-shell $4f^7$ and $4f^65d^1$ of Eu^{2+} , for instance, it returns to solve 33 462 linear equations with nine variables, leading to some misrepresentations of the parameters.⁵⁸ Therefore, we calculate the $F_k(\text{ff})$, $F_k(\text{fd})$ and $G_k(\text{fd})$ parameters from the radial wavefunctions R_{nl} of the 4f and 5d Kohn–Sham orbitals of the lanthanide ions following eqn (8)–(10), which is the subject of the next section.

The calculation of the F_k , G_k , and ζ_{nl} parameters

The importance of relativity in the physics of lanthanide elements is not negligible.^{59–63} There are different approaches dedicated to the implementation of relativistic corrections in DFT. Besides the spin-orbit interaction H_{SO} (eqn (11)), which has itself a relativistic origin; the physics behind the Dirac equation in quantum chemistry is reasonably well characterized by the scalar-relativistic equations.⁶⁴ We can perform scalar-relativistic calculations at the zeroth-order regular approximation (ZORA) to the Dirac equation^{65–69} or the first-order relativistic Pauli Hamiltonian^{70–74} in DFT. In Fig. 3, the solutions of the radial wavefunctions of the 4f and 5d Kohn–Sham orbitals obtained for a gaseous Eu^{2+} ion are graphically represented, where the influence of the relativistic correction is evaluated. A noticeable expansion of the R_{nl} is observed when relativistic corrections are implemented in the computational details (Fig. 3), in line with the definition of relativity acting on f and d orbitals.^{64,75} This expansion is severely pronounced for the Pauli-relativistic calculation (Fig. 3), because of the explicit insertion of the Darwin and mass-velocity terms in the master equation.^{70–74} The calculated $F_k(\text{ff})$, $F_k(\text{fd})$, $G_k(\text{fd})$ and ζ_{nl} parameters using R_{4f} and R_{5d} (Fig. 3) are collected in Table 2. In total, there are:³⁵

three $F_k(\text{ff})$ parameters: $F_2(\text{ff})$, $F_4(\text{ff})$ and $F_6(\text{ff})$; plus

two $F_k(\text{fd})$ parameters: $F_2(\text{fd})$ and $F_4(\text{fd})$; plus

three $G_k(\text{fd})$ parameters: $G_1(\text{fd})$, $G_3(\text{fd})$ and $G_5(\text{fd})$; plus

two spin-orbit coupling constants: ζ_{4f} and ζ_{5d} .³⁵

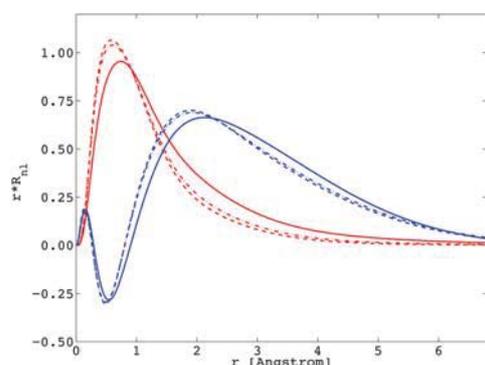


Fig. 3 Representation of the radial wavefunctions R_{nl} corresponding to the 4f (in red) and 5d (in blue) Kohn–Sham orbitals of a gaseous Eu^{2+} ion, obtained at the Pauli-relativistic (solid curve), the ZORA-relativistic (dotted-and-dashed curve) and the non-relativistic (dashed curve) levels of theory.

Table 2 Calculated Slater–Condon parameters and spin–orbit coupling constants (in cm^{-1}) obtained at the non-relativistic (a), the ZORA-relativistic (b) and the Pauli-relativistic (c) levels of theory, corresponding to the two-open-shell $4f^7$ and $4f^65d^1$ electron configurations of a gaseous Eu^{2+} ion

	Slater–Condon parameters and spin–orbit coupling constants		
	(a)	(b)	(c)
$F_2(\text{ff})$	500.19	475.60	388.47
$F_4(\text{ff})$	64.66	61.32	49.92
$F_6(\text{ff})$	6.87	6.51	5.30
$F_2(\text{fd})$	245.32	245.36	244.72
$F_4(\text{fd})$	17.86	18.12	18.82
$G_1(\text{fd})$	338.38	369.81	431.92
$G_3(\text{fd})$	29.97	31.66	35.34
$G_5(\text{fd})$	4.70	4.91	5.40
ζ_{4f}	2133.90	1980.90	1246.50
ζ_{5d}	1279.31	1245.93	987.25

The parameters in Table 2 were determined from the wavefunctions ψ_{4f} and ψ_{5d} , prepared in the same manner as was done for Fig. 1.

Note that the parameters $F_2(\text{ff})$, $F_4(\text{ff})$ and $F_6(\text{ff})$ are acting principally on the single-open-shell $4f^n$ configuration, but they are also present in the diagonal block of the $4f^{n-1}5d^1$ interaction matrix. Experimentally known spectral terms of the $4f^7$ configuration of Eu^{2+} concern only the ground state ^8S and the two excited states ^6P and ^6I ,⁷⁶ although there are 119 levels arising from the multi-electron configuration.⁷⁷ The calculated energy values of these ^8S , ^6P and ^6I spectral terms are given in Table 3, obtained using the parameters in Table 2. They are also compared with the available experimental data taken from the framework of the NIST atomic spectra database.⁷⁶

We determined the deviations between the calculated and the experimental spectral terms (Table 3) using eqn (16):

$$\varepsilon[\%] = 100 \times \frac{|E_{\text{calc}} - E_{\text{exp}}|}{E_{\text{exp}}} \quad (16)$$

For the three computational methods under consideration, we obtained a maximum deviation of 30.64%, 24.09% and 2.41% for the non-relativistic, ZORA-relativistic and Pauli-relativistic

Table 3 Calculated multiplet energy levels (calc.) of a gaseous Eu^{2+} ion (in cm^{-1}) at non-relativistic (a), ZORA-relativistic (b) and Pauli-relativistic (c) levels of theory, compared with the experimentally-known spectral terms (exp.) corresponding to the $4f^7$ electron configuration

	Calc.			Exp. ^a
	(a)	(b)	(c)	
$^8\text{S}_{7/2}$	0.00	0.00	0.00	0.00
$^6\text{P}_{7/2}$	36379.05	34596.14	28854.94	28200.06
$^6\text{P}_{5/2}$	37400.88	35526.15	29317.04	28628.54
$^6\text{P}_{3/2}$	38339.68	36381.98	29758.39	— ^a
$^6\text{I}_{7/2}$	40277.33	38282.44	31591.89	31745.99
$^6\text{I}_{9/2}$	40978.41	38917.65	31888.23	31954.21
$^6\text{I}_{17/2}$	41370.71	39274.11	32060.93	32073.30
$^6\text{I}_{11/2}$	41542.51	39430.07	32135.09	32179.55
$^6\text{I}_{15/2}$	41901.08	39756.07	32293.32	32307.78
$^6\text{I}_{13/2}$	41881.85	39739.21	32287.83	32314.14

^a Taken from ref. 76 where the energy value of $^6\text{P}_{3/2}$ is not known.

calculations, respectively. On the other hand, we also obtained a minimum deviation of 26.87%, 20.59% and 0.04%. There is an appropriate agreement between the Pauli-relativistic results and the experimental data. Experimentally-known spectral terms of the $4f^65d^1$ configuration of Eu^{2+} in ref. 76 assemble some states of octet and sextet spin multiplicity, which in tensor operator techniques represent the direct product: $4f^6(^7F) \otimes ^2D$, giving rise to the following terms: 8P , 6P , 8D , 6D , 8F , 6F , 8G , 6G , 8H and 6H . Note that the whole manifold of the $4f^65d^1$ configuration allows the consideration of 906 spectral terms, including not only the high octet spin multiplicity but also the lower sextet, quartet and doublet, the energies of which are obtained from the DFT calculation using the $F_k(\text{ff})$, $F_k(\text{fd})$, $G_k(\text{fd})$, ζ_{4f} and ζ_{5d} parameters (Table 2) and the $A(\text{fd})$ parameter discussed in the previous section.

The calculated deviations ε (eqn (16)) from the experimentally-known spectral terms⁷⁶ are represented in Fig. 4 for the three theoretical methods under consideration. Here also the Pauli-relativistic calculation leads to the best reproduction of the experimental data, its mean deviation being 6.17% (Fig. 4), which is far smaller than those obtained at the non-relativistic and ZORA-relativistic levels of theory.

In this section, the impact of the relativistic correction on the spectroscopy of lanthanide ions is clearly justified; an appropriate description of the radial R_{4f} and R_{5d} wavefunctions is a prerequisite, enabling a good reproduction of the experimental data.

Structural analysis of the doped systems

The doping of lanthanide ions into solid state materials is nowadays a topic of significant interest due to the optical effects.³⁷ There are several instrumental methods to probe the local structure around the impurity ions in solid state compounds, such as nuclear magnetic resonance (NMR),⁷⁸ extended X-ray absorption fine structure (EXAFS),^{79,80} as well as electron paramagnetic resonance (EPR).⁸¹ However, these methods do not give direct results of the local geometry, offering only data that can be corroborated to it. A clear answer is found in the theoretical side, mimicking the doping of solid state materials by means of

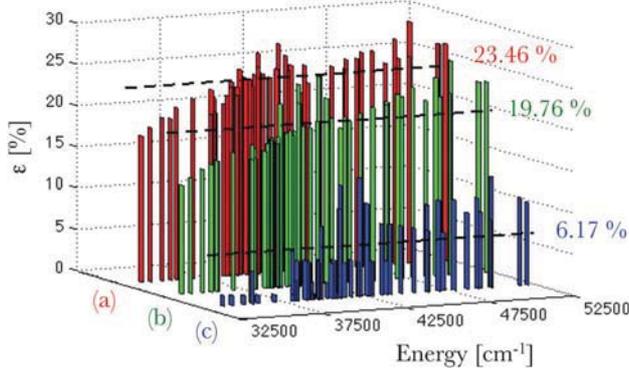


Fig. 4 Representation of the error distribution ε (in %) with respect to the experimental data⁷⁶ of the calculated multiplet energy levels corresponding to the $4f^65d^1$ configuration of a gaseous Eu^{2+} ion, at the non-relativistic (in red, (a)), ZORA-relativistic (in green, (b)) and Pauli-relativistic (in blue, (c)) levels of theory. The calculated mean deviations from the experimental data are also given.

band structure methods. In this section, we investigate the local structure around the Eu^{2+} impurity incorporated in the CaF_2 and SrCl_2 matrices. Calcium fluoride (CaF_2) and strontium chloride (SrCl_2) belong to the cubic $Fm\bar{3}m$ space group (No. 225).^{41,42} The divalent Eu^{2+} ion enters in the matrices in the site formally occupied by Ca^{2+} and Sr^{2+} . It is then coordinated by eight fluoride or chloride ligands, respectively, in the systems $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$, within the O_h point group.

For the pristine CaF_2 and SrCl_2 systems (Fig. 5), the calculated lattice parameters are given in Table 4 in terms of the DFT functional used in the band structure algorithm. It is found that both GGA and LDA calculations yield different lattice equilibrium constants (Table 4), *i.e.* different local relaxations. In terms of a direct comparison, we consider the GGA calculation most appropriate to simulate the experimental data, although the cells are slightly larger than the experimental ones.

For the $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$ systems, we constructed super-cells which double the number of the unit cells of CaF_2 and SrCl_2 in the a , b and c directions. The Eu^{2+} ion was placed in the position (0,0,0). In these cases, the super-cells are big enough inasmuch as the interactions between two Eu^{2+} ions are minimized. We relaxed the positions of the atoms, fixing the lattice parameters to the theoretical values obtained for the pure systems. This mimics the resistance of the whole lattice against defect-induced distortions, under the conditions of a lower doping concentration than the $2 \times 2 \times 2$ super-cells actually worked upon have. The optimized Eu-F and Eu-Cl bond lengths are 2.4732 Å and 3.0774 Å, respectively, which represent an elongation with respect to the Ca-F and Sr-Cl bond lengths obtained for the pure systems: 2.3893 Å and 3.0515 Å.

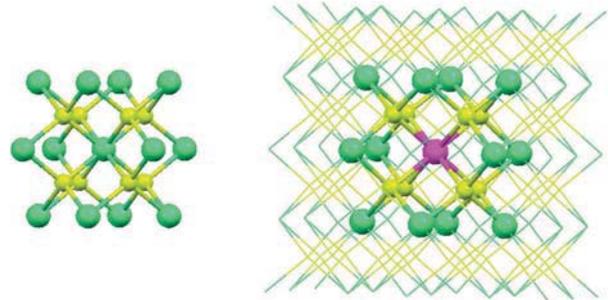


Fig. 5 Representation of the crystal structure of CaF_2 showing the unit cell (left hand side). The local structure of an Eu^{2+} centre embedded in a $2 \times 2 \times 2$ unit cell of CaF_2 (right hand side). Colour code: Ca^{2+} in green, F^- in yellow and Eu^{2+} in violet. For clarity, some Ca^{2+} and F^- ions are represented with a wireframe shape.

Table 4 Calculated lattice parameters a , b , c (in Å) and α , β , γ (in Å) obtained for CaF_2 and SrCl_2 crystallizing in the cubic $Fm\bar{3}m$ space group (No. 225), compared with experimental X-ray diffraction data

	CaF_2			SrCl_2		
	LDA	GGA	Exp. ^a	LDA	GGA	Exp. ^b
a , b , c	5.3342	5.5179	5.4355	6.8088	7.0472	6.965
α , β , γ	90.0	90.0	90.0	90.0	90.0	90.0

^a Taken from ref. 82. ^b Taken from ref. 83.

The description of the local structure of doped materials is important in the further evaluation of the ligand field Hamiltonian (eqn (13)), the presence of the impurity in the host materials producing distortions due to differences in the ionic radii or electronic structure. We favoured here the band structure algorithms for geometrical purposes, although we can certainly conceive of a cluster geometry optimization approach, which is already popular in computational chemistry, especially while dealing with excited states geometry.^{36,39}

The calculation of B_q^k ligand field parameters

In the $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$ systems, the site symmetry of the Eu^{2+} impurity is O_h and the non-zero Wybourne parameters are as follows:

$$B_0^4(f,f), B_4^4(f,f), B_{-4}^4(f,f), B_0^6(f,f), B_4^6(f,f) \text{ and } B_{-4}^6(f,f),$$

for the sub-matrix corresponding to the $\langle f|H_{\text{LF}}|f \rangle$ (eqn (13));

$$B_0^4(d,d), B_4^4(d,d) \text{ and } B_{-4}^4(d,d),$$

for the sub-matrix corresponding to the $\langle d|H_{\text{LF}}|d \rangle$ (eqn (13)).

The inversion center in the O_h point group allows vanishing of the elements of the sub-matrix $\langle f|H_{\text{LF}}|d \rangle$.⁴⁷

The ligand field energy schemes of the 4f and the 5d orbitals of Eu^{2+} in the $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$ systems were calculated taking the cubic clusters $(\text{EuF}_8)^{6-}$ and $(\text{EuCl}_8)^{6-}$, respectively, which have the optimized geometries obtained in the previous section. Point charges were placed at the coordinates of the next-neighbouring Ca^{2+} and Sr^{2+} ions, which are also shown as ball-and-sticks in the super-cell in Fig. 5. These were used in order to mimic the long-range interaction of the crystal hosts.

The ligand field energies and wavefunctions were obtained from Kohn–Sham orbitals of restricted DFT calculations within the AOC reference, by evenly placing six electrons in the 4f orbitals and one electron in the 5d. We previously presented the analysis of the ligand field interaction with respect to the change of the DFT functional for the two-open-shell 4f and 5d problem in Pr^{3+} .³⁵ It was found that, in the 5d ligand field, the DFT functional does not play an important role, whereas in the 4f, the hybrid B3LYP functional is required in order to obtain realistic ligand field parameters.³⁵ Therefore we have used B3LYP for the computation of the electronic structure of Eu^{2+} .

The 4f orbitals form the basis of t_{1u} , t_{2u} and a_{2u} irreducible representations (irreps) of the O_h point group. The 5d orbitals are in the basis of the e_g and the t_{2g} irreps. The values of the ligand field B_q^k parameters were determined by linear equation fitting using eqn (13), knowing the following ratios for the octahedral symmetry constraint:

$$B_4^4(l, l) = B_{-4}^4(l, l) = \sqrt{\frac{5}{14}} B_0^4(l, l), \quad (17)$$

with l standing for d and f, and

$$B_4^6(f, f) = B_{-4}^6(f, f) = -\sqrt{\frac{7}{2}} B_0^6(f, f). \quad (18)$$

The calculated values of the B_q^k parameters for the $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$ systems are presented in Table 5, together with

Table 5 Calculated ligand field parameters (calc.) in cm^{-1} obtained for the systems $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$, compared with available experimental data (exp.)

	$\text{CaF}_2:\text{Eu}^{2+}$		$\text{SrCl}_2:\text{Eu}^{2+}$	
	Calc.	Exp. ^a	Calc.	Exp. ^b
$B_0^4(f,f)$	-1765	-2386	-829	-1035
$B_4^4(f,f)$	-1055	-1430	-496	-619
$B_0^6(f,f)$	120	966	208	-761
$B_4^6(f,f)$	-225	-1807	-389	1423
$B_0^4(d,d)$	-34 821	-33 600	-21 086	-21 296
$B_4^4(d,d)$	-20 810	-20 080	-12 601	-12 727

^a Taken from ref. 41. ^b Taken from ref. 42.

the experimentally-deduced ones. For $B_q^k(d,d)$, the theoretical values are in good agreement with the experimental data.^{41,42} However, for $B_q^k(f,f)$, although the $B_0^4(f,f)$ and *ipso facto* the $B_4^4(f,f)$ (eqn (17)) are also in the magnitude of the experimental data, the $B_0^6(f,f)$ and related parameters (eqn (18)) are slightly underestimated for both the $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$ systems. This departure between the calculated and the experimental values is reflected primarily in the ordering of the 4f orbitals splitting. The ordering obtained from computation was as follows for both $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$ systems:

$$t_{1u} < t_{2u} < a_{2u},$$

while it resulted in the swapped sequence,

$$t_{1u} < a_{2u} < t_{2u},$$

for the experimentally-deduced parameters obtained for the $\text{SrCl}_2:\text{Eu}^{2+}$ system.⁴²

The change in the orbital ordering may be attributed to the impact of the neighbouring cations, where the symmetry-adapted linear combination of their virtual orbitals may stabilize the a_{2u} irrep. This is not achieved here in the small cluster models of $(\text{EuCl}_8)^{6-}$. Nevertheless, a direct comparison between $B_q^k(f,f)$ and $B_q^k(d,d)$ indicates that the effect of the 4f parameters will be completely superseded by the 5d ones.

AOM analysis of the ligand field interaction

For the sake of more intuitive insight, the B_q^k parameters can be converted to the AOM scheme,⁴⁷ reformulating the ligand field matrix in eqn (13) as follows:

$$\langle 3, m|H_{\text{LF}}|3, m' \rangle = \sum_{k=1}^{\text{ligands}} \sum_{\lambda=\sigma,\pi} D_{m,\lambda}^{4f}(k) \cdot D_{m',\lambda}^{4f}(k) \cdot e_{\lambda,k}(4f), \quad (19)$$

$$\langle 2, m|H_{\text{LF}}|2, m' \rangle = \sum_{k=1}^{\text{ligands}} \sum_{\lambda=\sigma,\pi} D_{m,\lambda}^{5d}(k) \cdot D_{m',\lambda}^{5d}(k) \cdot e_{\lambda,k}(5d), \quad (20)$$

where D^{4f} and D^{5d} are the matrix elements defined in terms of Euler angles (Wigner's Darstellungsmatrizen)^{6,7,47} and k is the running index for the ligand system. The $e_{\lambda} \equiv e_{\sigma}$, e_{π} parameters have the meaning of perturbation exerted by σ and π sub-components of

the density cloud of the ligands (or by corresponding overlap effects, in another heuristic formulation).

A general problem in establishing the parametric conversion is the fact that the AOM matrix is not traceless, the sum of the diagonal elements for a homoleptic $[ML_n]$ complex with linearly ligating ligands (isotropic π effects) being $n(e_\sigma + 2e_\pi)$, instead of zero, like in the standard ligand field model. In the case of the 4f shell, in octahedral symmetry, the situation does not impinge upon the parametric conversion since we have two independent parameters, $B_0^4(f,f)$ and $B_0^6(f,f)$ in the Wybourne scheme (Table 5), *versus* two AOM parameters $e_\sigma(4f)$ and $e_\pi(4f)$, uniquely related to the two relative gaps in the ligand field splitting in O_h symmetry.

The mutual conversion is done by the following formulas:

$$e_\sigma(4f) = -\frac{9}{44}B_0^4(f,f) + \frac{63}{1144}B_0^6(f,f); \quad (21)$$

$$e_\pi(4f) = -\frac{3}{22}B_0^4(f,f) - \frac{189}{1144}B_0^6(f,f). \quad (22)$$

The comparison of computed *versus* experimental fitted Wybourne parameters (Table 5) can be regarded as semi-quantitative in general, with certain apparent mismatches, as is the opposite sign found for the $B_0^6(f,f)$ value in the case of the $SrCl_2:Eu^{2+}$ system. The same sign mismatch in the case of the computed *versus* fitted $B_4^6(f,f)$ is not an independent feature, given the mentioned $B_4^6(f,f)/B_0^6(f,f)$ proportionality (eqn (18)). The conversion to AOM parameterization allows a certain assessment of the situation. Thus, the calculated AOM parameters for the $SrCl_2:Eu^{2+}$ system are as follows (in cm^{-1}): $e_\sigma(4f) = 181.02$ and $e_\pi(4f) = 78.68$, while the conversion of reported fitted $B_0^4(f,f)$ and $B_0^6(f,f)$ values⁴² yielded the following values: $e_\sigma(4f) = 169.79$ and $e_\pi(4f) = 266.86$. One observes that the experimental values lead to the somewhat counterintuitive situation of $e_\pi(4f) > e_\sigma(4f)$ values, with the result then that the numeric experiment may be, in a relative sense, a more reliable source, not for absolute values but with respect to the inter-parametric ratios. The fact is that the ligand field parameters on the 4f shell show small values, in general, being prone to fit uncertainties given the large number of active parameters. The reference work⁴² considered a fit with several empirical terms such as Trees and Marvin corrections, while keeping imposed fixed ratios among the more fundamental Slater–Condon parameters, and therefore the full comparability of computed *versus* fitted parameters is partly hindered, considering that we worked here only by first-principle leverages: ligand field, Slater–Condon and spin–orbit coupling parameters, without other degrees of freedom.

For the 5d shell, the single gap between e_g and t_{2g} does not need the two AOM parameters, so that must impose certain conventions, like the $e_\sigma(5d)/e_\pi(5d) = 3$ ratio.³⁵ However, we have not advanced in this direction, given the good match of the computed and fitted 5d-type B_q^k parameters, which do not demand the call of AOM as a further moderator in the comparative discussion.

The ligand field interaction, besides lifting the degeneracy of the 4f and 5d orbitals, also has a side effect expanding the

radial wavefunctions towards the ligands positions. This is commonly known as the nephelauxetic effect, a concept coined by C. K. Jørgensen⁸⁴ which is the subject of the next section.

The nephelauxetic effect

The nephelauxetic effect describes the fact that the parameter values of the inter-electron repulsion are usually smaller in complexes than in the corresponding free ions.⁸⁴ The word nephelauxetic was created by basic translation of “cloud expansion” from Greek. We can quantitatively analyze the changes in the metal wavefunctions with respect to the presence of ligands, underlying the action of the nephelauxetic effect. The 4f shell is shielded from the interaction with the chemical environment inasmuch as independently to the ligand type, the reduction of the free ion inter-electron repulsion $F_k(ff)$ parameters are negligible.^{47,85} On the other hand, the virtual 5d and 6s shells are able to interact with the neighbourhood, ensuring therefore the bonding of lanthanide ions.^{86–88} We present in Fig. 6 the radial wavefunction R_{5d} of Eu^{2+} in the presence of eight fluoride and eight chloride ligands in a cubic arrangement. For comparison purposes, we also present the radial wavefunction obtained for the gaseous Eu^{2+} free ion (Fig. 3). One observes the pronounced expansion of R_{5d} in the series of fluoride and chloride ligands, highlighting the overlap of ligands by the orbitals from the lanthanide ion. Due to this effect, as explained in previous instances, in the excited states of a $4f^{n-1}5d^1$ lanthanide configuration, the calculated bond lengths are always shorter than those obtained in the ground $4f^n$ configuration.^{36,39}

Recalling eqn (9) and (10), we calculated the $F_k(fd)$ and $G_k(fd)$ parameters in the complex, based on the radial shapes shown in Fig. 6. Compared with Fig. 3, one notes that R_{4f} remains almost the same, while R_{5d} were shifted by the nephelauxetic effect (see also ref. 47).

The results are given in Table 6, together with the calculated spin–orbit coupling constant ζ_{5d} , using eqn (12) and the $\Delta(fd)$ gap. All the parameters (Table 6) are reduced when compared to the Pauli-relativistic quantities in Table 2. The nephelauxetic ratio β is defined as the fraction made from the inter-electron

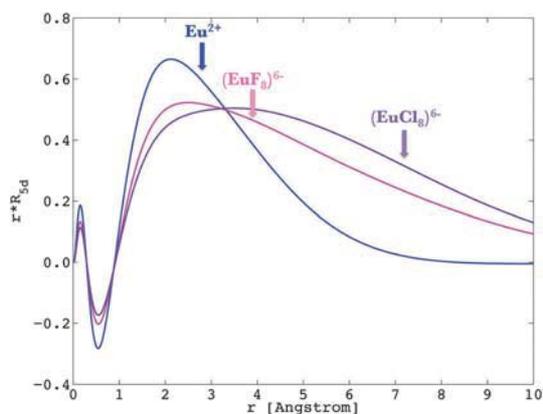


Fig. 6 Representation of R_{5d} of Eu^{2+} in the free ion (in blue), in $(EuF_8)^{6-}$ (in pink) and $(EuCl_8)^{6-}$ (in violet), obtained at the Pauli-relativistic level of theory.

Table 6 Calculated values of the Slater–Condon $F_k(\text{fd})$ and $G_k(\text{fd})$, the spin–orbit coupling ζ_{5d} and the $\Delta(\text{fd})$ gap (in cm^{-1}) obtained for the systems $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$, compared with the experimentally-deduced values

	$\text{CaF}_2:\text{Eu}^{2+}$			$\text{SrCl}_2:\text{Eu}^{2+}$		
	Calc.	β	Exp. ^a	Calc.	β	Exp. ^a
$F_2(\text{fd})$	138.42	0.57	133.33	100.56	0.41	117.43
$F_4(\text{fd})$	9.88	0.53	10.25	6.79	0.36	8.54
$G_1(\text{fd})$	232.08	0.54	192.29	160.56	0.37	162.06
$G_3(\text{fd})$	18.22	0.52	17.30	12.31	0.35	14.41
$G_5(\text{fd})$	2.74	0.51	2.72	1.84	0.34	2.26
ζ_{5d}	505.76	0.51	760	371.14	0.38	844
$\Delta(\text{fd})$	18 800	—	23 500	12 400	—	—

^a The $F^k(\text{fd})$ and $G^k(\text{fd})$ are taken from ref. 41 and 42. They are converted to the corresponding $F_k(\text{fd})$ and $G_k(\text{fd})$ parameters using the conversion factor in ref. 43.

parameters obtained in the complex and in the free ion, for instance:

$$\beta(F_2(\text{fd})) = \frac{F_2(\text{fd})_{\text{Complex}}}{F_2(\text{fd})_{\text{Free ion}}} \quad (23)$$

The calculated β values for $F_k(\text{fd})$, $G_k(\text{fd})$ and ζ_{5d} were also given in Table 5. We calculate mean β values of 0.53 and 0.37 for $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$, respectively.

We can obtain the $\Delta(\text{fd})$ gap for the $\text{CaF}_2:\text{Eu}^{2+}$ system from ref. 41 which we compare with our calculated value (Table 6). Unfortunately, the experimental value for the same parameter was not specified for the $\text{SrCl}_2:\text{Eu}^{2+}$ system.⁴² The difference between the calculated $\Delta(\text{fd})$ gap and that obtained in ref. 41 is directly related to the $F_k(\text{ff})$ parameters (Table 2), which is also present in the diagonal elements of the CI matrix of the $4f^65d^1$ configuration of Eu^{2+} . Since the values of our calculated $F_k(\text{ff})$ parameters are larger than that given in ref. 41 our $\Delta(\text{fd})$ is accordingly smaller.

The $4f^7-4f^65d^1$ transitions

The multiplet energy levels corresponding to the $4f^7$ and the $4f^65d^1$ configurations of Eu^{2+} doped into CaF_2 and SrCl_2 are given in Fig. 7 in the spectral range of 0–55 000 cm^{-1} (those for the entire spectral range are given in the ESI,† Fig. S1 and S2). They were computed in the LFDFT algorithm using the following non-empirical parameters: $F_k(\text{ff})$ and ζ_{4f} (Table 2); $F_k(\text{fd})$, $G_k(\text{fd})$, ζ_{5d} and $\Delta(\text{fd})$ (Table 6); and B_q^k 's parameters (Table 5).

The transitions from the initial $4f^7(^8S_{7/2})$ state to the final $4f^65d^1$ are electric dipole-allowed, with the calculation of the electric dipole transition moments obtained from eqn (15). The oscillator strength for the zero phonon lines between the ground state $^8S_{7/2}$ of $4f^7$ and the final states of $4f^65d^1$ are calculated and represented in Fig. 7. The most intense transitions are given with respect to the irreps of the octahedral double group. In the circumstances of a non-degenerate $^8S_{7/2}$ state of the $4f^7$ sub-system, the energies of the $4f^7-4f^65d^1$ transitions are practically the same as the position of $4f^65d^1$ spectral terms. The intensities were computed by corresponding handling of the dipole moment represented in the ligand field CI basis, depending all on a single reduced matrix element, ultimately irrelevant as an

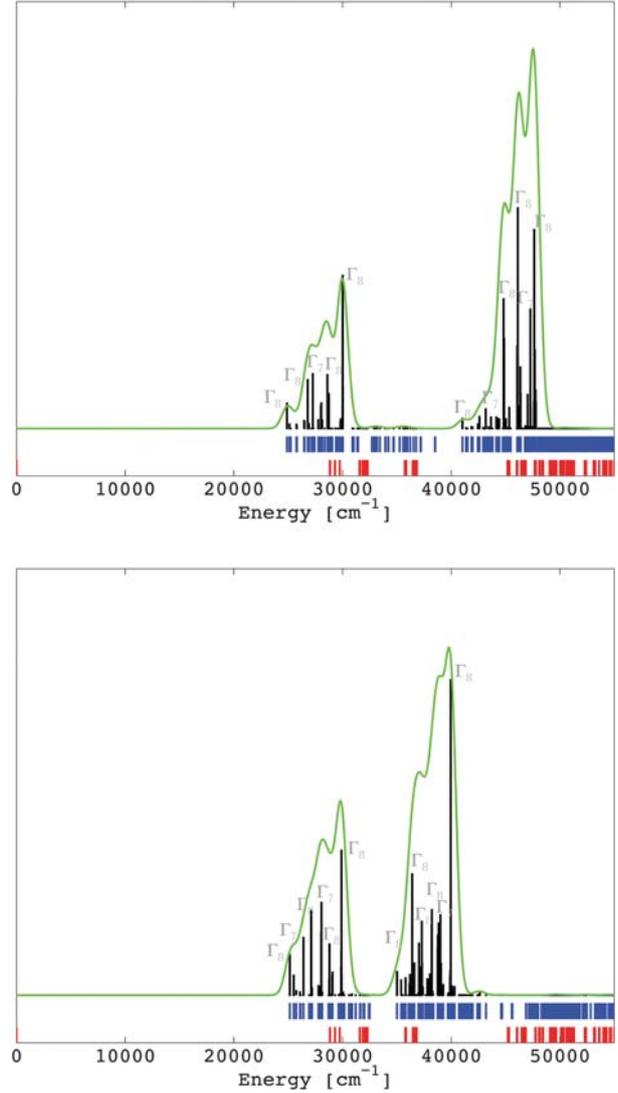


Fig. 7 Calculated multiplet energy levels from the $4f^7$ (in red) and $4f^65d^1$ (in blue) configurations of Eu^{2+} in CaF_2 (up) and SrCl_2 (down), together with the intensities of the excitation $4f^7 \rightarrow 4f^65d^1$ transitions, *i.e.* zero phonon lines (in black). The green curve represents a superimposition of a Gaussian band with a width of 500 cm^{-1} on the zero phonon lines.

absolute value, if we consider an arbitrary scale of spectral rendering. The zero-field splitting, which transforms the $^8S_{7/2}$ state of $4f^7$ to $\Gamma_6 + \Gamma_7 + \Gamma_8$ in the actual octahedral symmetry, is in the magnitude of tenths of cm^{-1} . The $4f^65d^1$ transitions are characterized by two dominant bands (Fig. 7), in line with the excitation spectrum seen in ref. 41 and 42 for $\text{CaF}_2:\text{Eu}^{2+}$ and $\text{SrCl}_2:\text{Eu}^{2+}$. The correspondence between the theoretical results and the excitation spectrum is seen in the ESI,†, where the excitation spectra of $\text{CaF}_2:\text{Eu}^{2+}$ (Fig. S1, ESI,†) and $\text{SrCl}_2:\text{Eu}^{2+}$ (Fig. S2, ESI,†) are reproduced from ref. 41 and 42.

Conclusions

Optical and magnetic effects in lanthanide-based compounds are phenomena intimately understood with the help of ligand

field theory. In this work, we have outlined important points for a realistic description of the electronic structure and the optical properties of Eu^{2+} -doped CaF_2 and SrCl_2 compounds. The treatment of the local distortions due to the presence of the Eu^{2+} impurity in the fluorite structures of CaF_2 and SrCl_2 was addressed by periodical band structure calculation. The LFDFT algorithm was used for the calculation of the multiplet energy levels of the $4f^7$ and $4f^65d^1$ electron configurations of Eu^{2+} . The optical $4f^7-4f^65d^1$ transitions were determined, with a good qualitative agreement achieved between the non-empirical investigations and the experimental findings. In particular, the convoluted calculated spectrum can be immediately compared with experimental data, thus showing the usefulness of the approach to experimental scientists. The computational methods and post-computational analyses contained in the LFDFT algorithm are producing reliable ligand field and related parameters, consolidating the academic insight into the structure–property relationships of rare-earth materials and paving the way to the *desiderata* of property engineering.

There are several advantageous characteristics that this fully non-empirical LFDFT method possesses which should be noted and remembered, besides the predictive capability, very important today for the vast number and kind of rare-earth-based technological materials. The method can be applied to any lanthanide ions, for general $4f^n-4f^{n-1}5d^1$ transitions with different coordination symmetries. The LFDFT approach has other advantages against widespread semi-empirical and full *ab initio* methods, not least the fact that it can be applied to larger-size systems in a relatively short computational time.

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Notes and references

- H. Bethe, *Ann. Phys.*, 1929, **295**, 133.
- J. H. Van Vleck, *J. Chem. Phys.*, 1935, **3**, 807.
- J. S. Griffith, *The Theory of Transition Metal Ions*, Cambridge University Press, Cambridge, 1961.
- M. Gerloch, J. H. Harding and R. G. Woodley, *Struct. Bonding*, 1981, **46**, 1.
- B. N. Figgis and M. A. Hitchman, *Ligand Field Theory and Its Applications*, Wiley-VCH, New York, 2000.
- C. Schäffer and C. Jørgensen, *Mol. Phys.*, 1965, **9**, 401.
- W. Urland, *Chem. Phys.*, 1976, **14**, 393.
- C. Daul, *Int. J. Quantum Chem.*, 1994, **52**, 867.
- C. Daul, E. J. Baerends and P. Vernooijs, *Inorg. Chem.*, 1994, **33**, 3538.
- M. Atanasov, C. A. Daul and C. Rauzy, *Struct. Bonding*, 2004, **106**, 97.
- L. Noodleman and J. G. Norman, *J. Chem. Phys.*, 1979, **70**, 4903.
- L. Noodleman, *J. Chem. Phys.*, 1981, **74**, 5737.
- M. Mitani, V. Mori, Y. Takano, D. Yamaki, Y. Yoshioka and K. Yamaguchi, *Chem. Phys.*, 2000, **113**, 4035.
- A. Bencini and F. Totti, *Int. J. Quantum Chem.*, 2005, **101**, 819.
- F. Cimpoesu, B. Frescus, C. I. Oprea, H. Ramanantoanina, W. Urland and C. Daul, *Mol. Phys.*, 2015, DOI: 10.1080/00268976.2015.1007107.
- M. Atanasov, C. Rauzy, P. Baettig and C. Daul, *Int. J. Quantum Chem.*, 2005, **102**, 119.
- M. Atanasov, C. Daul, H. U. Gudel, T. A. Wesolowski and M. Zbiri, *Inorg. Chem.*, 2005, **44**, 2954.
- D. Reinen, M. Atanasov and W. Massa, *Z. Anorg. Allg. Chem.*, 2006, **632**, 1375.
- M. Atanasov, P. Comba and C. Daul, *Inorg. Chem.*, 2008, **47**, 2449.
- M. Gruden-Pavlovic, M. Peric, M. Zlatar and P. Garcia-Fernandez, *Chem. Sci.*, 2014, **5**, 1453.
- Z. D. Matovic, M. S. Jeremic, R. M. Jelic, M. Zlatar and I. Z. Jakovljevic, *Polyhedron*, 2013, **55**, 131.
- G. Morrison, H. Ramanantoanina, W. Urland, M. D. Smith and H.-C. zur Loye, *Inorg. Chem.*, 2015, **54**, 5504–5511.
- C. A. Daul, *J. Phys.: Conf. Ser.*, 2013, **428**, 012023.
- A. Borel, C. A. Daul and L. Helm, *Chem. Phys. Lett.*, 2004, **383**, 584.
- F. Senn, L. Helm, A. Borel and C. A. Daul, *C. R. Chim.*, 2012, **15**, 250.
- M. Atanasov and C. A. Daul, *Chem. Phys. Lett.*, 2003, **379**, 209.
- M. Atanasov, P. Comba and C. Daul, *Inorg. Chem.*, 2006, **110**, 13332.
- I. Ciofini and C. Daul, *Coord. Chem. Rev.*, 2003, **238–239**, 187.
- M. Atanasov and C. A. Daul, *C. R. Chim.*, 2005, **8**, 1421.
- M. Atanasov, E. J. Baerends, P. Bruyndonckx, C. Daul, C. Rauzy and M. Zbiri, *Chem. Phys. Lett.*, 2004, **399**, 433.
- F. Senn and C. A. Daul, *THEOCHEM*, 2010, **954**, 105.
- F. Senn, M. Zlatar, M. Gruden-Pavlovica and C. Daul, *Monatsh. Chem.*, 2011, **142**, 593.
- L. Petit, A. Borel, C. Daul, P. Maldivi and C. Adamo, *Inorg. Chem.*, 2006, **45**, 7382.
- M. Atanasov and C. Daul, *Chimia*, 2005, **59**, 504.
- H. Ramanantoanina, W. Urland, F. Cimpoesu and C. Daul, *Phys. Chem. Chem. Phys.*, 2013, **15**, 13902.
- H. Ramanantoanina, W. Urland, A. Garcia-Fuente, F. Cimpoesu and C. Daul, *Chem. Phys. Lett.*, 2013, **588**, 260.
- S. Nakamura and G. Fasol, *The Blue Laser Diode*, Springer, Berlin, 1997.
- H. Ramanantoanina, W. Urland, A. Garcia-Fuente, F. Cimpoesu and C. Daul, *Phys. Chem. Chem. Phys.*, 2014, **16**, 14625.
- H. Ramanantoanina, W. Urland, B. Herden, F. Cimpoesu and C. Daul, *Phys. Chem. Chem. Phys.*, 2015, **17**, 9116.
- B. Herden, A. Garcia-Fuente, H. Ramanantoanina, T. Jüstel, C. Daul and W. Urland, *Chem. Phys. Lett.*, 2015, **620**, 29.
- G. W. Burdick, A. Burdick, V. Deev, C.-K. Duan and M. F. Reid, *J. Lumin.*, 2006, **118**, 205.
- Z. Pan, L. Ning, B.-M. Cheng and P. A. Tanner, *Chem. Phys. Lett.*, 2006, **428**, 78.

- 43 R. D. Cowan, *The theory of atomic structure and spectra*, University of California Press, Berkeley, 1981.
- 44 H. Lueken, *Magnetochemie*, Teubner Studienbücher Chemie, Stuttgart, 1999.
- 45 M. Gerloch, *Magnetism, ligand-field analysis*, Cambridge University Press, Cambridge, 1983.
- 46 S. Hübner, *Optical spectra of transparent rare earth compounds*, Academic Press, New York, 1978.
- 47 H. Ramanantoanina, W. Urland, F. Cimpoesu and C. Daul, *Phys. Chem. Chem. Phys.*, 2014, **16**, 12282.
- 48 G. te Velde, F. M. Bickelhaupt, S. J. A. van Gisbergen, C. F. Guerra, E. J. Baerends, J. G. Snijders and T. Ziegler, *J. Comput. Chem.*, 2001, **22**, 931.
- 49 C. F. Guerra, J. G. Snijders, G. te Velde and E. J. Baerends, *Theor. Chem. Acc.*, 1998, **99**, 391.
- 50 E. J. Baerends, T. Ziegler, J. Autschbach, D. Bashford, A. Berces, F. M. Bickelhaupt, C. Bo, P. M. Boerrigter, L. Cavallo, D. P. Chong, L. Deng, R. M. Dickson, D. E. Ellis, M. van Faassen, L. Fan, T. H. Fischer, C. F. Guerra, A. Ghysels, A. Giammona, S. J. A. van Gisbergen, A. W. Götz, J. A. Groeneveld, O. V. Gritsenko, M. Grüning, S. Gusarov, F. E. Harris, P. van den Hoek, C. R. Jacob, H. Jacobsen, L. Jensen, J. W. Kaminski, G. van Kessel, F. Kooststra, A. Kovalenko, M. V. Krykunov, E. van Lenthe, D. A. McCormack, A. Michalak, M. Mitoraj, J. Neugebauer, V. P. Nicu, L. Noodleman, V. P. Osinga, S. Patchkovskii, P. H. T. Philipsen, D. Post, C. C. Pye, W. Ravenek, J. I. Rodriguez, P. Ros, P. R. T. Shipper, G. Schreckenbach, J. S. Seldenthuis, M. Seth, J. G. Snijders, M. Sola, M. Swart, D. Swerhone, G. te Velde, P. Vernooijs, L. Versluis, L. Vissler, O. Visser, F. Wang, T. Wesolowski, E. M. van Wezenbeek, G. Wiesenekker, S. K. Wolff, T. K. Woo, A. L. Yarkolev, *ADF2013.01*, available at <http://www.scm.com>.
- 51 P. J. Stephens, F. J. Devlin, C. F. Chabalowski and M. J. Frisch, *J. Phys. Chem.*, 1994, **98**, 11623.
- 52 (a) G. Kresse and J. Hafner, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1993, **47**, 558; (b) G. Kresse and J. Furthmüller, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169.
- 53 S. H. Vosko, L. Wilk and M. Nussair, *Can. J. Phys.*, 1980, **58**, 1200.
- 54 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865.
- 55 P. E. Blöchl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953.
- 56 G. Kresse and D. Joubert, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, 1758.
- 57 T. Ziegler, A. Rauk and E. J. Baerends, *Theor. Chim. Acta*, 1977, **43**, 261.
- 58 See also the two different approaches devoted to the calculation of the Slater–Condon parameters in the two-open-shell $4f^2$ and $4f^15d^1$ configurations of Pr^{3+} in ref. 35 and 39.
- 59 P. Pyykkö and J. P. Desclaux, *Acc. Chem. Res.*, 1979, **12**, 276.
- 60 K. S. Pitzer, *Acc. Chem. Res.*, 1979, **12**, 271.
- 61 C. Clavaguera, J.-P. Dognon and P. Pyykkö, *Chem. Phys. Lett.*, 2006, **429**, 8.
- 62 P. Pyykkö, *Chem. Rev.*, 1988, **88**, 563.
- 63 P. Pyykkö, *Chem. Rev.*, 2012, **112**, 371.
- 64 D. Cremer, W. Zou and M. Filatov, *WIREs Comput. Mol. Sci.*, 2014, **4**, 436.
- 65 E. van Lenthe, A. E. Ehlers and E. J. Baerends, *J. Chem. Phys.*, 1999, **110**, 8943.
- 66 E. van Lenthe, E. J. Baerends and J. G. Snijders, *J. Chem. Phys.*, 1993, **99**, 4597.
- 67 E. van Lenthe, E. J. Baerends and J. G. Snijders, *J. Chem. Phys.*, 1993, **101**, 9783.
- 68 E. van Lenthe, J. G. Snijders and E. J. Baerends, *J. Chem. Phys.*, 1996, **105**, 6505.
- 69 E. van Lenthe, R. van Leeuwen, E. J. Baerends and J. G. Snijders, *Int. J. Quantum Chem.*, 1996, **57**, 281.
- 70 J. G. Snijders and E. J. Baerends, *Mol. Phys.*, 1978, **36**, 1789.
- 71 J. G. Snijders, E. J. Baerends and P. Ros, *Mol. Phys.*, 1979, **38**, 1909.
- 72 T. Ziegler, J. G. Snijders and E. J. Baerends, *J. Chem. Phys.*, 1981, **74**, 1271.
- 73 T. Ziegler, V. Tschinke, E. J. Baerends, J. G. Snijders and W. Ravenek, *J. Phys. Chem.*, 1989, **93**, 3050.
- 74 P. M. Boerrigter, M. A. Buijse and J. G. Snijders, *Chem. Phys.*, 1988, **122**, 357.
- 75 R. D. Cowan and D. C. Griffin, *J. Opt. Soc. Am.*, 1976, **66**, 1010.
- 76 A. Kramida, Y. Ralchenko, J. Reader and NIST ASD Team, *NIST Atomic Spectra Database (ver. 5.2)*, [Online], National Institute of Standards and Technology, Gaithersburg, MD, 2014, available: <http://physics.nist.gov/asd>, April 21, 2015.
- 77 C. W. Nielson and G. F. Koster, *Spectroscopic Coefficients for the p^n , d^n and f^n Configurations*, The MIT Press Cambridge, 1963.
- 78 R. E. Walstedt and L. R. Walker, *Phys. Rev. B: Solid State*, 1974, **9**, 4857.
- 79 P. A. O'Day, J. J. Rehr, S. I. Zabinsky and G. E. Brown, *J. Am. Chem. Soc.*, 1994, **116**, 2938.
- 80 P. Ghina, A. Carollo, G. Flor, L. Malavasi and G. S. Peruga, *J. Phys. Chem. B*, 2005, **109**, 4365.
- 81 B. V. Padlyak, M. Grinberg, T. Lukasiewicz, J. Kisielewski and M. Swirkowicz, *J. Alloys Compd.*, 2003, **361**, 6.
- 82 D. N. Batchelder and R. O. Simmons, *J. Chem. Phys.*, 1964, **41**, 2324.
- 83 H. Ott, *Z. Kristallogr.*, 1926, **63**, 222.
- 84 C. K. Jørgensen, *Absorption Spectra and Chemical Bonding in Complexes*, Pergamon Press, Oxford, 1962.
- 85 P. A. Tanner and Y. Y. Yeung, *J. Phys. Chem. A*, 2013, **117**, 10726.
- 86 D. L. Clark, J. C. Gordon, P. J. Hay and R. Poli, *Organometallics*, 2005, **24**, 5747.
- 87 F. Cimpoesu, N. Dragoie, H. Ramanantoanina, W. Urland and C. Daul, *Phys. Chem. Chem. Phys.*, 2014, **16**, 11337.
- 88 M. L. Neidig, D. L. Clark and R. L. Martin, *Coord. Chem. Rev.*, 2013, **257**, 394.